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VSEPR Theory
Basic
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VSEPR and Molecular

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Geometry: Rules, Examples, and Practice<del>How To</del> <del>Draw Lewis</del>

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Chemistry #24
VSEPR Theory and
Molecular
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Practice Problem: VSEPR Theory and Molecular Geometry Chemistry -Molecular Structure (2.5 of 59) VSEPR Notation (Watch Page 8/51

this version instead) <del>VSEPR</del> Theory: Memorising Tip to learn Various Shapes in Vsepr Theory (Best Shortcut) VSEPR Theory Lewis Dot Structure Practice Problems (with answers and Page 9/51

explanation)
Valence Bond
Theory, Hybrid
Orbitals, and
Molecular
Orbital Theory
Lewis Dot
Structures VSEPR
Theory

VSEPR Theory +
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Introduction, Formal Charge, Molecular Geometry, Resonance, Polar or Nonpolar Electron Domains, VSEPR and Determining Molecular Geometries 12. The Shapes of Molecules: VSEPR Theory Chapter 9 Page 13/51

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Extension VSEPR Worksheet
#2 For this
activity, you
will need to
refer to your
Lewis Dot
Worksheet # 1.

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Molecular Geometry Activity Answers Author: i;½i;½le arncabq.ctsnet.o rg-Nadine Eberha rdt-2020-08-28-0 5-44-56 Subject: i;½i;½Chemistry Molecular Geometry Activity Answers

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school chemistry students. Be sure to download the lab sheet below before you begin. Molecular Shape and the VSEPR Theory Lab Sheets.

Molecular Geometry Chemistry Worksheet Page 21/51

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Molecular Geometry Explain the difference between a bonding electron domain and a nonbonding electron domain using the examples in Model 1. 10. Circle the correct word or Page 24/51

phrase to complete the sentences: Pairs of electrons will (attract/repel

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are not all flat structures. Each has a threedimensional shape that helps account for its various chemical and physical properties. [ 9-12 Content Standard B -Structure and properties of matterl Students Page 29/51

often find it difficult to

Activity Answers The molecular geometry main shapes are tetrahedral, trigonal planar, trigonal pyramidal, bent, Page 30/51

and linear and are named by measuring the bond angles between the central atom and another atom bonded to it.

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Geometry
Worksheet & Lab
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In this activity you will learn how to predict molecular shapes. Model 1 Lewis Structures Lone pair =  $\cdot \cdot$  H 2 CO 3 electron domains (3 bonding, 0 nonbonding) 3-D Molecular Shape BeF 2 2 electron Page 32/51

domains (2 bonding, 0 nonbonding) CH 4 4 electron domains (4 bonding, 0 nonbonding) NH 3 4 electron domains (3 bonding, 1 nonbonding) H 2 O 4 electron domains

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electron pair geometry: molecular geometry: shape: polarity: BeH 2: 2: 0: linear: MX 2: linear: nonpolar: BH 3: 0: trigonal

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planar: MX 3: trigonal planar: nonpolar: SO 2: 2.91 trigonal planar: MX 2 E: bent: polar : CH tetrahedral: MX 4: tetrahedral: nonpolar: NH 3: 3: 1: tetrahedral : MX 3 E: trigonal . pyramidal. Page 35/51

polar: H 2 0: 2: 2: tetrahedral: MX 2 E 2: bent or V-shaped. polar: PCl 5: 5: 0: trigonal . bipyramidal. MX 5: trigonal . bipyramidal. nonpolar: SF 6: 6: 0

Molecular
Modeling
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Activity // How to Determine Molecular Geometry -YouTube: This video describes one method for quickly finding the major geometrical shapes for simple molecules. Molecular Page 37/51

Geometries The VSEPR theory describes five main shapes of simple molecules: linear, trigonal planar, tetrahedral, trigonal bipyramidal, and octahedral.

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<del>keyemis</del>trv Building Molecules Molecular Geometry VERSION (Advanced Assignment) Students will be able to .... · Draw the Lewis Dot Structure for molecules. Page 50/51

Identify the number of Bonded Atoms and Lone Pairs on a central atom. Identify the bond angle in molecules.

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